

Introduction

Reaction between cations and anions will be examined in this lab. The resulting precipitate is indicative of the reaction that occurred.

In our lab a lead nitrate solution, $\text{Pb}(\text{NO}_3)_2$ will be mixed with a solution of potassium chromate, K_2CrO_4 , a yellow precipitate will form. This precipitate must be a combination of K^+ or Pb^{2+} cation and CrO_4^{2-} or NO_3^{2-} anion. Since most potassium salts and nitrates are soluble in water, PbCrO_4 must be the precipitate. The equation for this reaction is:

