

IA *The Alkali Metals*

In order of increasing atomic number the members of group IA of the periodic table are:

hydrogen,
lithium,
sodium,
potassium,
rubidium,
cesium and
francium.

With the exception of hydrogen, which is a gas and which frequently imparts a quality of acidity to its compounds, the other members of the group display rather striking similarities of chemical behaviour, all reactive with H_2O to form strongly alkaline solutions. The elements in the group, including hydrogen, are characterized by a valence of one, having one electron in an outer shell available for reaction. Because of their chemical similarities, these elements, along with ammonium and sometimes magnesium, are considered the sixth group in classical qualitative chemical analysis separations.